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a.) C_5H_{12} . b.) $C_{10}H_{22}$. c.) C_2H_2 . 45. Ethanol
(CH_3CH_2OH) and propane (C_3H_8) have similar
molar masses. Does that mean they have the same

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boiling point? Explain. 46. For each of the pairs below, circle the substance with the higher boiling point: a.) I₂ and F₂. b.) HF and LiF. c.) CH₂O and CCl₄. 47.

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C 12.1 Pure substances and mixtures AQA GCSE
Chemistry C12 Chemical Analysis Kerboodle Answers:
Page No. 181. 1. Answer. In advertisement, a product added nothing to it is called as pure. For example: a pure orange juice without additives. While in chemistry- A pure substance is one that is made up of just one substance i.e either element or ...

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New(9-1) AQA GCSE Chemistry C12 Chemical Analysis

...

Chapter 12 Part B1: Additional Alkyl and Aryl Radicals
2 points Draw the structural or semi-structural formulas for all the alkyl/aryl radicals of the following. Place a X in place of the hydrogen which is removed to make the radical: (Not in Chapter 12-use the Internet)

Chapter 12 Homework Packet - fscj.me

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... 1.1 Assessed Homework Task (mark scheme) 1.1

Test (mark scheme) More Exam Questions on 1.1

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Atomic Structure and 1.4 Periodicity (mark scheme)

1.1 Atomic Structure - A-Level Chemistry

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Chapter 12 – Introduction to Organic Chemistry:

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Hydrocarbons Section 12.1 – Alkanes Goal: Identify properties characteristic of organic or inorganic compounds. Summary: Organic compounds: always contain carbon and hydrogen (though sometimes other nonmetals as well.) Inorganic compounds: all other compounds. Often ionic (metal + nonmetal)

Chapter 12 Introduction to Organic Chemistry:
Hydrocarbons

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Chapter 12, Section 12.1 - 12.6. General Chemistry
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Ions . Section 3.1 (page 95) Questions 2-10.

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(12) disulfide bridges; salt bridges; H-bonding;

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dispersion forces; (metal-ion coordination is the 5th)
(13) Cysteine (Cys) is the amino acid that can form disulfide bridges. (14) The acidic amino acids, aspartic acid (Asp) & glutamic acid (Glu) and the basic amino acids, lysine (Lys), arginine (Arg), and histidine (His) can form salt bridges because their side chains are ionized.

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