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Solutions

These parameters depend on the sizes of the solute and solvent molecules and on how strongly the latter interact with each other. This suggests two important limiting cases for reactions in solution. For water at room temperature, k_1 is typically $10^9 - 10^{10} \text{ dm}^{-3} \text{ mol}^{-1} \text{ s}^{-1}$ and k_2 is around $10^{-9} - 10^{-10} \text{ dm}^{-3} \text{ mol}^{-1} \text{ s}^{-1}$.

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Solutions Reactions in aqueous solution in which $E_a > 20 \text{ kJ/mol}$ are likely to fall into this category.

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Alternatively, if the activation energy of the A+B reaction dominates the kinetics, and the reaction is

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LibreTexts 4 CHAPTER 1 The Basics of
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Chemical kinetics is the study of
chemical processes and rates of
reactions. This includes the analysis of
conditions that affect speed of a
chemical reaction, understanding
reaction mechanisms and transition
states, and forming mathematical
models to predict and describe a
chemical reaction. The rate of a
chemical reaction usually has units of
 sec^{-1} , however, kinetics experiments
may span several minutes, hours, or
even days.

Understand Chemical Kinetics and
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