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describes the bonding,  
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coordination complexes.

It represents an application of molecular orbital theory to

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complexes. A transition metal ion has nine

valence atomic orbitals - consisting of five nd,

one s, and three p

orbitals. These orbitals are of appropriate

energy to form bonding interaction with ligands.

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Ligand field theory, in chemistry, one of several theories that describe the electronic structure of coordination or complex compounds, notably transition metal complexes, which consist of a central metal atom surrounded by a group of electron-rich atoms or molecules called ligands. The

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## Ligand-Field Theory

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metal complexes. A third model, based on molecular orbital theory, was therefore developed that is known as ligand-field theory.

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by a group of electron-  
rich atoms or molecules  
called ligands. Page  
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(LFT) which is an extension of crystal field theory (CFT) developed in the 1930s by two giants, Bethe and Van Vleck.

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According to the Crystal Field Theory, ligands that have high spin are considered "weak field" and ligands that have low spin are considered "strong field." This relates to the colors seen in a coordination complex.

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The ligand field theory is a firm background to foresee the magnetic properties of metallic complexes  $ML_n$  (M, transition metal ion; L, molecule or ligand). The ligand field splitting ? oct between the energies of  $t_{2g}$  and  $e_g$  orbitals of an octahedral complex  $ML_6$  is shown in Fig. 4 .

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important characteristics  
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depends upon

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on the metals and the  
ligand donor orbitals.

*An Introduction to  
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In coordination  
chemistry, a ligand is an  
ion or molecule that  
binds to a central metal  
atom to form a  
coordination complex.

The bonding with the  
metal generally involves

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formal donation of one or more of the ligand's electron pairs. The nature of metal–ligand bonding can range from covalent to ionic. Furthermore, the metal–ligand bond order can range from one to three. Ligands are viewed as Lewis bases, although rare cases are known to involve Lewis acidic "ligands". Metals



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